

0040-4039(94)01233-4

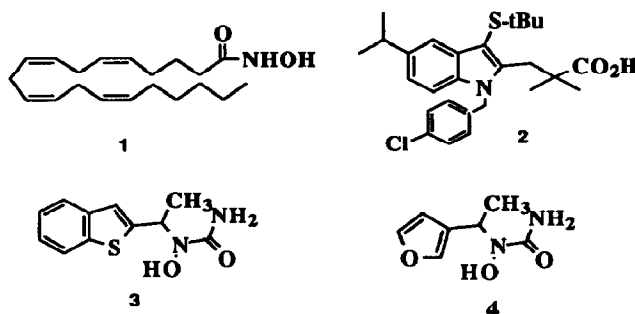
**SYNTHESIS OF SUBSTITUTED HETEROCYCLES. SIMPLE METHOD FOR THE INTRODUCTION OF THE N-HYDROXYUREA FUNCTIONALITY**

Yi-Yin Ku\*, Ramesh R. Patel, Brian A. Roden and David P. Sawick

Chemical and Agricultural Products Division  
 Abbott Laboratories, North Chicago, IL 60064-4000

**Abstract:** A simple method for the introduction of the N-hydroxyurea functionality has been developed. This involves conjugate addition of hydroxylamine to a vinylsulfone substrate. This method was successfully used in a facile synthesis of the 5-lipoxygenase inhibitors, zileuton and A-69412.

A variety of 5-lipoxygenase inhibitors have been the research focus of several groups. Examples of these compounds are arachidonate hydroxamic acid **1**,<sup>1</sup> MK-886 **2**,<sup>2</sup> zileuton **3** and A-69412 **4**. Zileuton and A-69412 contain an N-hydroxyurea as a necessary component for biological activity.<sup>3</sup> Our research efforts have been directed toward an efficient synthesis of **3** and **4**.

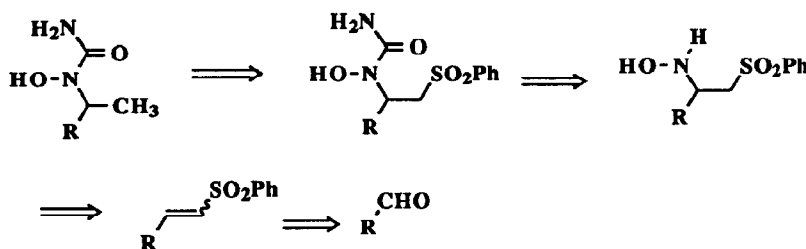


The most common method for preparation of an N-hydroxyurea is by treatment of an N-hydroxylamine with potassium isocyanate in the presence of acid. The hydroxylamine can be prepared by one of several methods: (1) acid catalyzed reduction of an oxime;<sup>4</sup> (2) addition of an organolithium compound to an oxime<sup>5</sup>

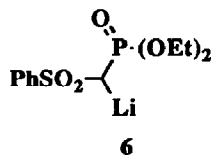
in the presence of  $\text{BF}_3\text{OEt}_2$ ; 6 or (3) displacement of a leaving group with hydroxylamine. All of these methods suffer from a variety of drawbacks. Method (1) is highly dependant on the nature of the reducing agent whereas methods (2) and (3) yield product in low yields.

For the synthesis of 4, method (1), utilizing pyridine borane, proved the easiest route, initially. Unfortunately, over 5 equivalents were needed to drive the reaction to completion. Other reducing methods were investigated but none yielded satisfactory results. It was also noted that the hydroxylamine was highly soluble in water which made isolation tedious and usually led to lower yields. Neither method (2) nor (3) were practical in our hands. These problems forced the development of a new synthetic approach for the hydroxylamine. The basis for our new protocol centered on the use of an appropriately substituted vinyl sulfone which could accept a functional group pre-disposed for conversion to an N-hydroxyurea (Scheme 1).

Scheme 1



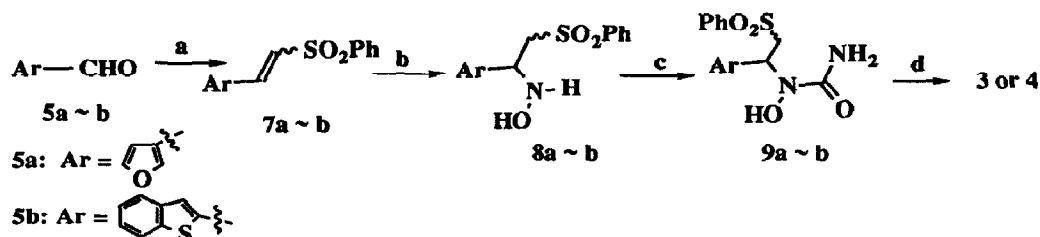
Scheme 2 details the synthesis. The vinyl sulfones were easily prepared via a Wittig-Horner olefin synthesis by reaction of the appropriate aldehyde (5a or b) with reagent 6<sup>7</sup> to give the vinyl sulfone derivatives 7.<sup>8</sup> Once the vinyl sulfones were in hand, the hydroxylamines were prepared by conjugate addition of either hydroxylamine or an O-protected hydroxylamine. For example, 7a was reacted with hydroxylamine



hydrochloride in the presence of  $\text{K}_2\text{CO}_3$  in THF to give 8a. Without further purification, this hydroxylamine was reacted with potassium isocyanate to give 9a. 7b was reacted with O-trimethylsilyl hydroxylamine to give 8b which was converted to 9b after reaction with trimethylsilyl isocyanate.

Removal of the sulfonyl group in the case of **9a** was easily achieved by simple treatment with magnesium in methanol to give **3**. In the case of **9b**, desulfonylation was not straight forward. Selective removal of the sulfone without decomposition was difficult, presumably through destruction of the benzothiophene core. A variety of desulfonylation methods were investigated ( $\text{Mg}^0/\text{MeOH}$ ;  $^9 \text{Na/Hg}$ ,  $\text{MeOH}$ ,  $\text{NaH}_2\text{PO}_4$ ;  $^{10} \text{NaHTe}$ ,  $\text{DMF}$  or  $\text{NaHTe}$ ,  $\text{EtOH}$ ;  $^{11} \text{SmI}_2$ ,  $\text{THF}$ ;  $^{12} \text{Na}_2\text{S}_2\text{O}_4$ ,  $\text{NaHCO}_3$ ,  $\text{DMF}$ ;  $^{13} \text{Zn}$ ,  $\text{HOAc}$ ;  $^{14} \text{Al/Hg}$ ,  $\text{MeOH}$ ,  $\text{NaH}_2\text{PO}_4$ ;  $^{15} \text{Ra-Ni}$ ,  $\text{EtOH}$ ;  $^{16} \text{Bu}_3\text{SnH}$ ,  $\text{AIBN}$ ,  $\text{Toluene}$ <sup>17</sup>). Of these methods, it was found that desulfonylation occurred most readily with buffered  $\text{Na/Hg}$  ( $\text{Na}_2\text{HPO}_4$ ,  $\text{THF}$ ,  $\text{H}_2\text{O}$ )<sup>10</sup> to give **3**.

### Scheme 2



Reagent: a) **6**,  $\text{LDA}$ ,  $\text{THF}$ ,  $-78^\circ\text{C}$  for **5a** to **7a** (80%), for **5b** to **7b** (83%); b)  $\text{NH}_2\text{OH}\cdot\text{HCl}$ ,  $\text{K}_2\text{CO}_3$ ,  $\text{THF}$ ,  $40^\circ\text{C}$  for **7a** to **8a** (82%),  $\text{TMSNHOH}$ ,  $\text{THF}$ , r.t. for **7b** to **8b** (85%); c)  $\text{KCNO}$ ,  $\text{EtOAc}$ ,  $\text{HCl}$ , r.t. for **8b** to **9b** (79%) or  $\text{TMSNCO}$ ,  $\text{THF}$ , r.t. for **8b** to **9b** (85%); d)  $\text{Mg/MeOH}$ , reflux for **3h** for **9a** to **4** (87%), 2.2%  $\text{Na/Hg}$ ,  $\text{Na}_2\text{HPO}_4$ ,  $\text{MeOH}$  for **9b** to **3** (81%)

We have presented here a simple method for the introduction of the N-hydroxylamine moiety to a heterocyclic system. The vinylsulfone substrate is readily available from the corresponding aldehyde and can be easily converted to its N-hydroxylamine derivative. Once converted to the N-hydroxyurea, desulfonylation liberates the desired product. This method is currently being investigated as a general method for the introduction of the N-hydroxyurea functionality to other interesting substrates.

### References and Notes

1. Corey, E.J.; Cashman, J.R.; Kantner, S.S.; Wright, S.W.; *J. Am. Chem. Soc.* **1984**, *106*, 1503.
2. Gillard, J.; Ford-Hutchinson, A.W.; Chan, C.; Charleson, S.; Denis, D.; Foster, A.; Fortin, R.; Leger, S.; McFarlane, C.S.; Morton, H.; Piechuta, H.; Reindeau, D.; Rouzer, C.A.; Rokach, J.; Young, R.; MacIntyre, D.E.; Peterson, L.; Bach, T.; Eiermann, G.; Hopple, S.;

- Humes, J.; Hupe, T.; Luell, S.; Metzger, J.; Meuer, R.; Miller, D.K. *Can. J. Physiol. Pharmacol.* **1989**, *67*, 456.
3. Carter, G.W.; Young, P.R.; Albert, D.H.; Bouska, J.; Dyer, R.; Bell, R.L.; Summers, J.B.; and Brooks, D.W. *J. Pharmacol. and Exper. Ther.* **1991**, *256*(3), 929.
  4. Kawase, M.; Kikugawa, Y. *J. Chem. Soc., Perkin Trans.* **1979**, *1*, 643.
  5. Richey, H.G. Jr.; McLane, R. C.; and Phillips, C. J. *Tetrahedron Lett.* **1976**, 233
  6. Rodriques, K. E.; Basha, A.; Summers, J. B.; and Brooks, D. W. *Tetrahedron Lett.* **1988**, *29*, 3455
  7. Reagent **6** prepared as follows:



- This is an extrapolation of the Wittig-Horner modification of the Wittig olefin synthesis described in: Camins, D.L.; Jacobine, A.F.; Marshall, J.L.; Turnbull, M.M. *Synthesis* **1978**, 309. For more information on the preparation and application of vinyl sulfones, see: Simpkind, N.S. *Tetrahedron*, **1990**, *46*, 6951.
8. Although the <sup>1</sup>H NMR of compound **6** indicated the presence of a single olefin isomer, no attempt was made to discern the exact olefin geometry.
  9. Brown, A.C.; Carpino, L.A. *J. Org. Chem.*, **1985**, *50*, 1749.
  10. Norman, B.H.; Gareau, Y., Padwa, A. *J. Org. Chem.*, **1991**, *56*, 2154.
  11. Huang, X.; Pi, J.-H. *Synth. Commun.*, **1990**, *20*, 2297.
  12. Kende, A.S.; Mendoza, J.S. *Tetrahedron Lett.*, **1990**, *31*, 7105.
  13. a) Harris, A.R.; Mason, T.J.; Hannah, G.R. *J. Chem. Res.*, **1990**, *5*, 218.  
b) Julia, M.; Launay, M.; Stoacino, J.P. *Tetrahedron Lett.*, **1982**, *23*, 2465.
  14. Huang, X.; Zhang, H.-Z. *Synth. Commun.*, **1989**, *19*, 97.
  15. Aluminum amalgam was prepared by stirring strips of aluminum foil in a 2% aqueous solution of mercuric chloride for @1 minute followed by washing in water then methanol. The amalgam was added directly to a methanolic solution of the sulfone buffered with @5 equivalents of Na<sub>2</sub>HPO<sub>4</sub>. Amalgam was added until tlc indicated complete loss of starting material.
  16. Trost, B.M.; Arndt, H.C.; Strege, P.E.; Verhoeven, T.R. *Tetrahedron Lett.* **1976**, 3478.
  17. Yoda, H.; Shirakawa, K.; Takabe, K., *Chemistry Letters*, **1989**, 1391

(Received in USA 10 August 1993; revised 27 May 1994; accepted 24 June 1994)